Tables d'enthalpie standard molaire de formation de corps composés à 25°c et 101300 Pa. Données en Kj/mole

COMPOSE	ETAT	$\Delta H°_f$
ALUMINIUM	s	0
$AlCl_3$	s	-704
	aq	-1033
Al_2O_3	s	-1670
ARGENT	s	0
AgCl	s	-127
$AgNO_3$	s	-124
	aq	-102
AZOTE	g	0
NH_3	g	-46
	aq	-80
NH_4Cl	s	-314
	aq	-300
NH_4NO_3	s	-365
	aq	-340
NO	g	+90
NO_2	g	+33
N_2O_4	g	+9
BARYUM	s	0
$Ba(OH)_2$	s	-945
$BaSO_4$	s	-1473
BROME	liq	0
	g	+31
	aq	-3
CALCIUM	s	0
$CaCO_3$	s	-1207
$CaCl_2$	s	-796
	aq	-877
CaO	s	-636
$Ca(OH)_2$	s	-986
$CaSO_4$	s	-1434

COMPOSE	ETAT	$\Delta H°_f$
CARBONE	s	0
CO	g	-111
CO_2	g	-393
	aq	-413
CHLORE	g	0
	aq	-23
HCl	g	-92
	aq	-167
CHROME	s	0
Cr_2O_3	s	-1140
COBALT	s	0
$CoCl_2$	s	-313
	aq	-392
$CoCl_2.6H_2O$	s	-2115
CUIVRE	s	0
$CuSO_4$	s	-771
	aq	-844
$CuSO_4.5H_2O$	s	-2280
FER	s	0
$FeCl_2$	aq	-423
$FeCl_3$	aq	-824
Fe_2O_3	s	-550
$FeSO_4$	s	-928
	aq	-998
HYDROGENE	g	0
H_2O	liq	-286
	g	-242
H_2O_2	liq	-188

COMPOSE	ETAT	$\Delta H°_f$
IODE	s	0
	g	+62
	aq	+23
HI	g	+26
MERCURE	liq	0
	g	+61
HgO(rouge)	s	-91
OXYGENE	g	0
PHOSPHORE		
blanc	s	0
rouge	s	-18
P_4O_{10}	s	-3042
H_3PO_4	aq	-1277
PLOMB	s	0
$PbCl_2$	s	-359
PbI_2	s	-175
PbO	s	-215
$PbSO_4$	s	-920
POTASSIUM	s	0
KBr	s	-394
	aq	-374
KCl	s	-437
	aq	-420
KI	s	-328
	aq	-306
KNO_3	s	-495
	aq	-460
K_2SO_4	s	-1434
	aq	-1409

COMPOSE	ETAT	$\Delta H°_f$
SODIUM	s	0
	g	+107
NaBr	s	-360
	aq	-362
Na_2CO_3	s	-1131
NaCl	s	-411
	aq	-407
NaI	s	-288
	aq	-295
$NaNO_3$	s	-468
	aq	-447
Na_2O	s	-414
Na_2O_2	s	-511
NaOH	s	-427
	aq	-469
Na_2S	s	-365
	aq	-447
Na_2SO_4	s	-1387
	aq	-1390
SOUFRE	s	0
	g	+279
H_2S	g	-21
	aq	-40
SO_2	g	-297
SO_3	g	-396
H_2SO_4	liq	-814
	aq	-909
ZINC	s	0
$ZnCl_2$	s	-415
	aq	-488